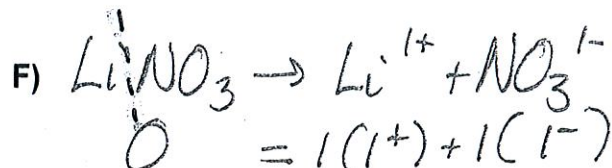
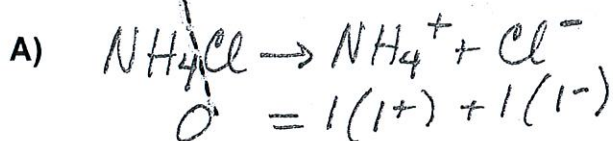


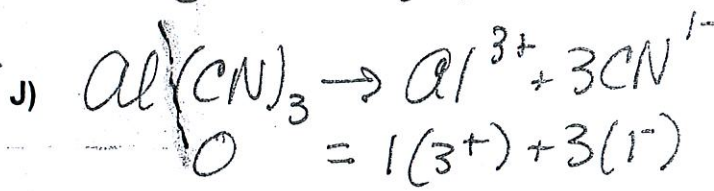
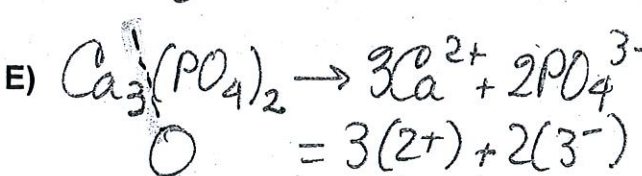
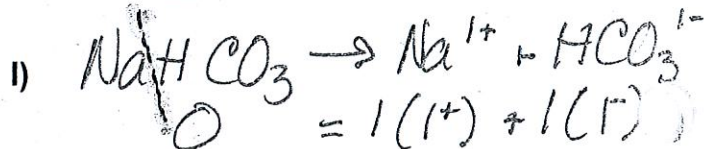
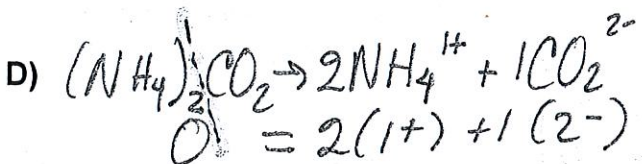
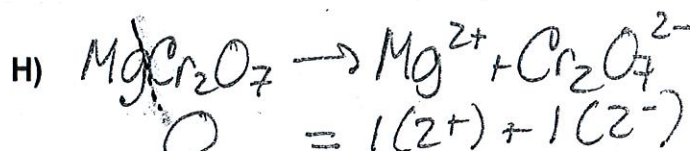
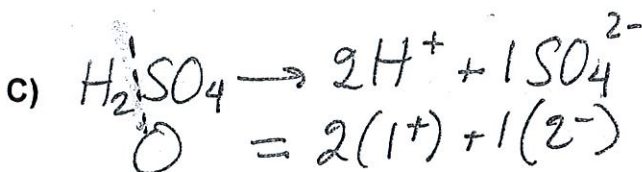
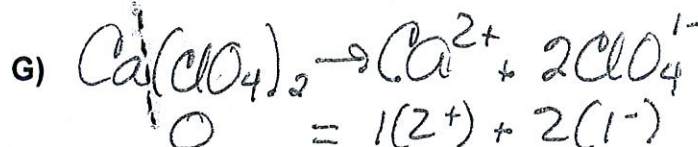
8. Quelle est la charge?

- A) Des métaux alcalins ? 1⁺ D) Des halogènes ? 1⁻
 B) Des alcalino-terreux ? 2⁺ E) Du radical OH ? 1⁻
 C) De l'oxygène ? 2⁻ F) Du radical NH₄ ? 1⁺

9. Écris l'équation de dissociation des composés du numéro 7 et justifie les charges de tes ions par le balancement des charges:



B) _____



10. En te basant sur les formules moléculaires, classe les substances suivantes selon qu'elles sont des **électrolytes(E)** ou des **non électrolytes(NE)**.

A) HNO₃ E

F) PCl₃ NE

B) C₃H₈ NE

G) CO₂ NE

C) NaCl E

H) KNO₃ E

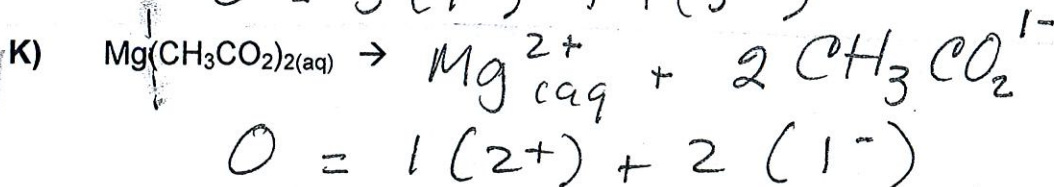
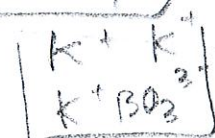
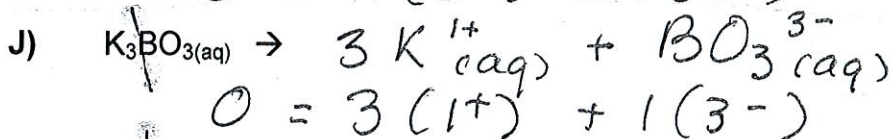
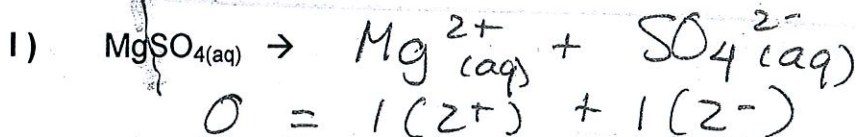
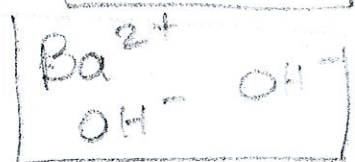
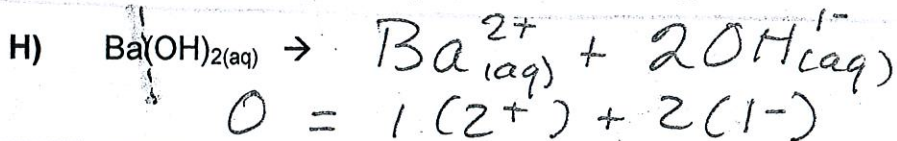
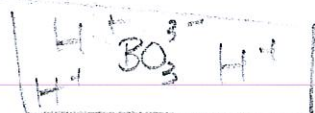
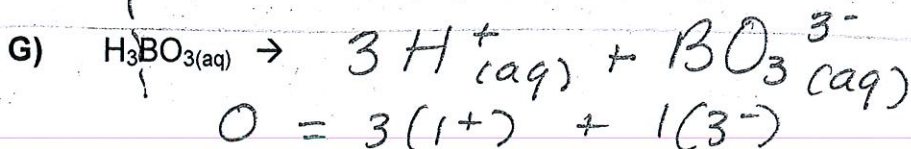
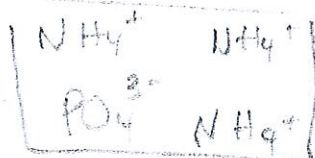
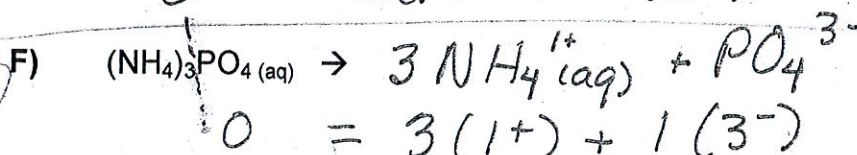
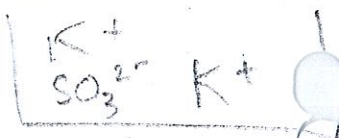
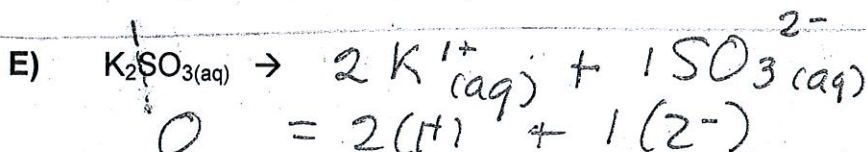
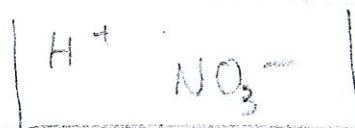
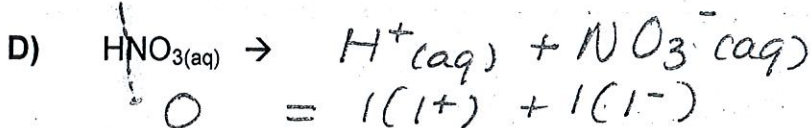
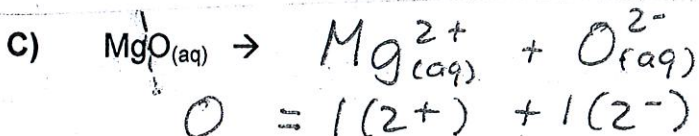
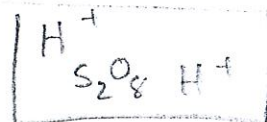
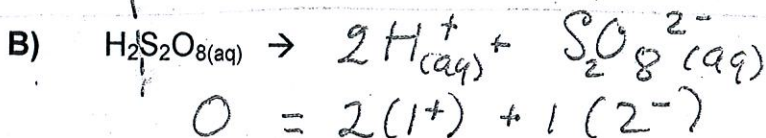
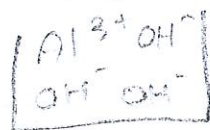
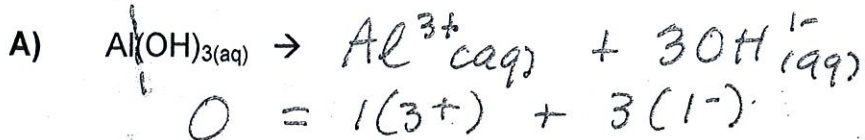
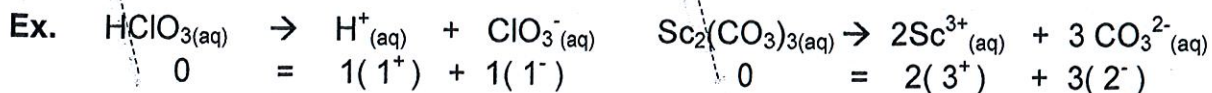
D) HBr E

I) LiOH E

E) NH₄OH E

J) H₃PO₄ E

11. Écris les équations de dissociation et justifie par le balancement des charges:
Indique par une ligne hachurée, l'endroit où la molécule se sépare.



* On "coupe" toujours après le métal ou le H ou le NH₄ (ou avant OH pour une base)